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(54) Title: PROCESS FOR THE CARBONYLATION OF ETHYLENICALLY OR ACETYLENICALLY UNSATURATED COM-
POUNDS

(57) Abstract: A process for the carbonylation of unsaturated compounds by contacting the unsaturated compound with carbon monoxide in the presence of a catalyst system comprising: (a) a source of palladium and/or platinum; and (e) an unsymmetrical bidentate diphosphine ligand of formula (I), $R^1R^2 > P^1 - R^3 - P^2 < R^4R^5$ (I) wherein P^1 and P^2 represent phosphorus atoms; R^3 represents a divalent organic bridging group; and R^1 , R^2 , R^4 and R^5 each individually, or R^1 and R^2 jointly, and/or R^4 and R^5 jointly represent organic groups that are covalently linked to the phosphorus; and wherein R^1 , R^2 , R^4 and R^5 are chosen in such way, that the phosphino group $R^1R^2 > P^1$ differs from the phosphino group $P^2 < R^4R^5$.



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PROCESS FOR THE CARBONYLATION OF ETHYLENICALLY OR
ACETYLENICALLY UNSATURATED COMPOUNDS

Field of the invention

The present invention provides a process for the carbonylation of ethylenically or acetylenically unsaturated compounds.

5 Background of the invention

In this specification, the term carbonylation refers to the reaction of alkenes or alkynes, i.e. ethylenically or acetylenically non-conjugated unsaturated compounds, under catalysis by a transition metal complex with carbon monoxide and a co-reactant having a mobile hydrogen atom, as for instance described in EP-A-0,495,548 and EP-A-0,565,199, WO-A-96/19434, WO-A-98/42717, WO-A-01/68583 and WO-A-01/72697. This process results in carboxylic acid products, or carboxylic acid derivatives, such as esters, amides or anhydrides, depending on the nature of the co-reactant.

EP-A-0,495,548 discloses a catalyst system comprising a palladium cation and diphosphine bidentate ligands bearing di-alkyl-substituted phosphino groups, such as for instance 1,3 bis(di-tertiary butyl phosphino)propane, and the use of such catalysts in various carbonylation reactions.

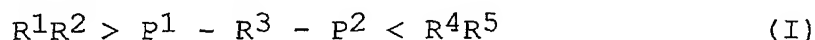
Although the disclosed catalyst systems show a good activity in carbonylation reactions, application of the process on an industrial scale requires further improvements in catalyst activity and selectivity for the desired products.

Summary of the invention

It has now been found that catalyst systems based on a metal cation selected from 8, 9 or 10 of the Periodic Table of Elements and an unsymmetrical diphosphine ligand exhibit improved catalyst activity in the carbonylation of ethylenically or acetylenically unsaturated compounds, thereby resulting in a significantly improved carbonylation process.

Accordingly, the subject invention relates to a process for the carbonylation of an ethylenically or acetylenically unsaturated compound by contacting the unsaturated compound with carbon monoxide and a co-reactant having a mobile hydrogen atom in the presence of a catalyst system comprising:

- (a) a source of a palladium and/or platinum; and
- (b) an unsymmetrical diphosphine of formula I,



wherein P^1 and P^2 represent phosphorus atoms; R^3 represents a divalent organic bridging group; and R^1 , R^2 , R^4 and R^5 each individually, or R^1 and R^2 jointly, and/or R^4 and R^5 jointly represent organic substituents that are covalently bound to the phosphorus atom; with the proviso that R^1 , R^2 , R^4 and R^5 are selected in such way, that the phosphino group $R^1R^2 > P^1$ is different from the phosphino group $P^2 < R^4R^5$.

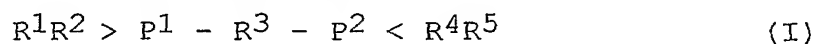
Detailed description of the invention

The subject process makes use of a catalyst system obtainable by combining (a) a source of a metal cation of palladium or, platinum with a diphosphine ligand as defined above. A source of one or more these metal cations may be used in the various carbonylation reactions. Of these metal cations, palladium (Pd) is

particularly preferred for carbonylation reactions involving olefinically unsaturated substrates, whereas platinum (Pt) is highly preferred for acetylenically unsaturated substrates, and for ethylenically unsaturated compounds if branched products are desired.

Examples of suitable metal sources are metal compounds such as salts of Pd and/or Pt and nitric acid, sulphuric acid or sulphonic acids, salts of Pd and Pt and/or carboxylic acids with up to 12 carbon atoms, Pd and/or Pt complexes, e.g. with carbon monoxide and/or acetyl acetate, or Pd and/or Pt combined with a solid material such as an ion exchanger. Pd and/or Pt acetate and acetyl acetate are examples of preferred metal sources.

The catalyst system further comprises as component (b) an unsymmetrical diphosphine of formula I,



wherein P^1 and P^2 represent phosphorus atoms, R^3 represents a divalent organic bridging group, and R^1 , R^2 , R^4 and R^5 each individually, or R^1 and R^2 jointly and/or R^4 and R^5 jointly represent organic substituents that are covalently bound to the phosphorus atom; with the proviso that R^1 , R^2 , R^4 and R^5 are selected in such way, that the phosphino group $R^1R^2 > P^1$ is different from the phosphino group $P^2 < R^4R^5$.

The beneficial effect of an unsymmetrical diphosphine compared to the corresponding symmetrical diphosphines is exhibited even if the substituents R^1 and R^2 attached to P^1 are very similar to the substituents R^4 and R^5 attached to P^2 . Without wishing to be bound to any particular theory, it is believed that the difference in substituents results in a difference in the electron

density at the phosphorus atoms. The presence of two binding phosphorus sites with different electronic density in the ligands then appears to have a beneficial effect on the reaction rate. The difference in electron donating ability of the phosphino fragments $P^1R^1R^2$ and $P^2R^4R^5$ of the subject ligands may conveniently be measured by method developed by C. Tolman for monodentate phosphine ligands as described in Chemical Reviews, 1977, Vol. 77, pages 313-348. In this method, the electron-donating ability of a phosphine ligand is determined by reacting one equivalent of a (monodentate) phosphine with $Ni(CO)_4$ to make $P^1R^1R^2$ a $Ni(CO)_3$ (phosphine) complex, and then measuring the carbonyl nCO IR stretching frequency (very sharp at high energy mode) of the $Ni(CO)_3$ (phosphine) complex. The more electron density the phosphine ligand donates to a metal centre, and hence the higher the electron-donating ability, the lower the nCO IR stretching frequency measured. In the diphosphine (b), even if only one of the substituents R^1 , R^2 , R^4 or R^5 differs from the other substituents, the two phosphorus atoms will have a different electron density.

The substituents R^1 to R^4 are thus selected in such way, that one of the phosphorus atoms P^1 or P^2 has a lower electron density than the other phosphorus atom, as illustrated by a separate, lower nCO IR stretching frequency that can be measured for a transition metal carbonyl complex of the ligand (b). The difference in electron density between the phosphorus atoms P^1 or P^2 may conveniently be achieved by selecting at R^1 , R^2 , R^4 and R^5 in such way, that one of the phosphorus atoms P^1

or P² has at least one more electron donating substituent R¹ and/or R², as compared to R⁴ and R⁵, or vice versa.

Particularly high catalyst activities were found with ligands wherein R¹ and R² together, or each individually
5 represent electron-donating groups as compared to a PPh₂-group, whereas R⁴ and R⁵ together or each individually represent an electron-accepting group as compared to a PPh₂-group.

Within the context of the subject specification, the
10 term "organic group" represents an unsubstituted or substituted, aliphatic or arylaliphatic radical having from 1 to 30 carbon atoms, which is covalently connected to the phosphorus atom by a carbon atom.

The organic groups R¹, R², R⁴ and R⁵ may each
15 independently be a monovalent group, or R¹ and R² together and/or R⁵ and R⁶ together may be divalent groups. The groups may further contain one or more heteroatoms such as oxygen, nitrogen, sulfur or phosphorus and/or be substituted by one or more
20 functional groups comprising for example oxygen, nitrogen, sulfur and/or halogen, for example by fluorine, chlorine, bromine, iodine and/or a cyano group.

Although increased catalyst activity as compared to each of the analogous symmetrical ligands was shown by
25 all non-symmetric bidentate ligands tested, a particularly high increase in catalyst activity was found with ligands wherein at least one of R¹ and R², or R⁴ and R⁵ represents an organic group containing a tertiary carbon atom through which this group is linked to the
30 phosphorus atom.

A yet higher activity was found where both R¹ and R², or R⁴ and R⁵ represent an organic group containing a tertiary carbon atom through which each group is linked to the phosphorus atom, and an even higher and yet
5 unprecedented activity was found where all of R¹, R², R⁴ and R⁵ represent an organic group containing a tertiary carbon atom through which each group is linked to the phosphorus atom.

Therefore, most preferably, the subject invention
10 also relates to the process for the carbonylation wherein all of R¹, R², R⁴ and R⁵ represent an organic group containing a tertiary carbon atom through which each group is linked to the phosphorus atom.

The tertiary carbon atom through which one or more of
15 the groups R¹, R², R⁴ or R⁵ preferably are connected to the phosphorus atom tertiary carbon atom, i.e. a carbon atom covalently bonded to the phosphorus and to three substituents other than hydrogen, may be substituted with aliphatic, cycloaliphatic, or aromatic substituents, or
20 may form part of a substituted saturated or non-saturated aliphatic ring structure, all of which may contain heteroatoms.

Preferably, the tertiary carbon atom of R¹ and R², or R⁴ and R⁵ may be substituted with alkyl groups, thereby
25 making the tertiary carbon atom part of a tertiary alkyl group, or by ether groups. Examples of suitable organic groups are tertiary-butyl, 2-(2-methyl)butyl, 2-(2-ethyl)butyl, 2-(2-phenyl) butyl, 2-(2-methyl) pentyl, 2-(2-ethyl)pentyl, 2-(2-methyl-4-phenyl)pentyl, 1-(1-methyl) cyclohexyl groups and 1-adamantyl-groups (wherein
30 1 or 2 adamantyl or trioxa-adamantyl-groups are bonded to phosphorus atom via a tertiary bridge head atom).

Although the groups R^1 and R^2 , or R^4 and R^5 may be each individually different organic groups, due to the use of lower amounts of different raw materials in the synthesis the groups R^1 and R^2 , or R^4 and R^5 preferably represent the same tertiary organic group. Yet more preferably, R^1 and R^2 , or R^4 and R^5 represent tertiary butyl groups. Accordingly, the subject invention also pertains to the process, wherein either both R^1 and R^2 , or both R^4 and R^5 represent a tertiary butyl group.

The substituents of the second phosphorus atom, i.e. R^1 and R^2 together, or R^4 and R^5 together preferably represent a divalent group that is directly attached to the phosphorus atom via two tertiary carbon atoms. This divalent group may have a monocyclic or a polycyclic structure.

Accordingly, R^1 and R^2 together or R^4 and R^5 together may also represent an optionally substituted divalent cycloaliphatic group, wherein the cycloaliphatic group is linked to the phosphorus atom via two tertiary carbon atoms. Examples of preferred divalent groups are unsubstituted or substituted C_4 - C_{30} -alkylene groups in which CH_2 - groups may be replaced by ether linkages through oxygen atoms, or other hetero groups. Suitable divalent groups include 1,1,4, 4-tetramethyl-buta-1,4-diyl-, 1,4-dimethyl-1, 4-dimethoxy-buta-1,4-diyl-, 1,1,5,5-tetramethyl-penta-1, 5-diyl-, 1,5-dimethyl-1,5-dimethoxy-penta-1,5-diyl-, 3-oxa-1,5-dimethoxy-penta-1, 5-diyl-, 3-oxa-1,1,5, 5-tetramethyl-penta-1,5-diyl-, 3-oxa-1,5-dimethyl-1, 5-dimethoxy-penta-1,5-diyl- and similar divalent radicals.

Particularly preferred monocyclic structures comprising R¹ and R² together, or R⁴ and R⁵ together include 2,2,6, 6-tetrasubstituted phosphinan-4-one or -4-thione structures. Ligands comprising such structures may be conveniently obtained under mild conditions as described in Welcher and Day, Journal of Organic Chemistry, J. Am. Chem. Soc., 27 (1962) 1824-1827, for instance by reaction of a secondary phosphine with 2,6-dimethyl-2,5-heptadien-4-one (also known as diisopropylidene acetone, or phorone).

A particularly preferred polycyclic structure including R¹ and R² or R⁴ and R⁵ together is for instance the 2-phospha-tricyclo[3.3.1.1{3,7}]decyl group that is substituted in 1,3 and 5 position (thus providing the tertiary carbon atoms through which the group is connected to the phosphorous atom), or a derivative thereof in which one or more of the carbon atoms are replaced by heteroatoms.

Tricyclo[3.3.1.1{3,7}]decane is the systematic name for a compound more generally known as adamantane. The 1,3,5-trisubstituted 2-phospha-tricyclo[3.3.1.1{3,7}]decyl group or a derivative thereof will thus be referred to as "2-PA" group (as in 2-phosphadamantyl group) throughout the specification. The 2-PA group is substituted on one or more of the 1, 3, 5 positions, and optionally also on the 7 position, with a monovalent organic group R⁷ from 1 to 20 atoms, preferably from 1 to 10 carbon atoms, yet more preferably from 1 to 6 carbon atoms. Examples of R⁷ include methyl, ethyl, propyl and phenyl.

More preferably, the 2-PA group is substituted on each of the 1, 3, 5 and 7 positions, suitably with identical groups R⁷, yet more preferably with methyl

groups. The 2-PA group further contains preferably additional heteroatoms other than the 2-phosphorus atom in its skeleton. Suitable heteroatoms are oxygen and sulphur atoms. More suitably, these heteroatoms are found in the 6, 9 and 10 positions. The most preferred bivalent radical is the 2-phospha-1,3,5,7-tetramethyl-6,9,10-trioxadamantyl group.

Very good results have been obtained in the carbonylation of ethylenically unsaturated compounds with catalyst system comprising bidentate diphosphine ligands whereby R^1 and R^2 are each individually organic groups only connected to each other via a phosphorus atom, whereas R^5 and R^6 together represent a bivalent organic group that is bonded to the second phosphorus atom via two tertiary carbon atoms.

Accordingly, especially preferred diphosphine ligands (b) for the catalyst system according to the subject invention are compounds according to formula (II), wherein R^1 together with R^2 , or R^4 together with R^5 , together with the respective phosphorus atoms P^1 or P^2 , respectively, form a 2-phospha-1, 3,5,7-tetramethyl-6,9,10-trioxadamantyl group, or a 2,2,6, 6-tetramethyl phosphinan-4-one, and wherein the other substituents R^1 and R^2 , or R^4 and R^5 each individually represent a tertiary butyl group.

Within the subject application, by a "divalent organic bridging group" R^3 is understood the group connecting the phosphorus atoms P^1 and P^2 via the shortest connection in the ligand. The bridging group R^3 preferably may have a structure as represented by formula (III),



wherein R^a and R^b independently represent the same or different optionally substituted methylene groups; R^b represents an organic group comprising the bivalent
5 bridging group C^1-C^2 through which R^b is connected to R^a and R^C ; m and n independently represent a natural number in the range of from 0 to 3, wherein the rotation about the bond between the carbon atoms of the bridging group C^1 and C^2 of the bridging group is restricted at a
10 temperature in the range of from 0 °C to 250 °C, and wherein the dihedral angle between the plane occupied by the three atom sequence composed of C^1 , C^2 and the atom directly bonded to C^1 in the direction of P^1 , and the plane occupied by the three atom sequence C^1 , C^2 and the
15 atom directly bonded to C^2 in the direction of P^2 , is in the range of from 0 to 120°. The terms bond and rotation are as defined in Hendrickson, Cram and Hammond, Organic
Chemistry, 3rd Edition, 1970, pages 175 to 201. Rotation
20 according to the subject invention means that the atoms attached to C^1 and C^2 respectively rotate about the axis that runs through the centre of the bond between C^1 and C^2 . The rotation about a bond is called "free" when the rotational barrier is so low that different conformations are not perceptible as different chemical species on the
25 time scale of the experiment. The inhibition of rotation of groups about a bond due to the presence of a sufficiently large rotational barrier to make the phenomenon observable on the time scale of the experiment is termed "hindered rotation" or "restricted rotation"
30 (as defined in IUPAC Compendium of Chemical Terminology,

2nd Edition (1997), 68, 2209). A suitable experiment can for instance be an ¹H-NMR-experiment as described in Hendrickson, Cram and Hammond, Organic Chemistry, 3rd Edition, 1970, pages 265 to 281 and in F.A. Bovey, 5 Nuclear Magnetic Resonance Spectroscopy, (New York, Academic Press, 1969), p. 1-20, provided that there are hydrogen atoms present in the ligand that will exhibit a suitable shift influenced by the bond between C¹ and C².

Preferably, there is no free rotation about the bond 10 between C¹ and C² at the temperature range at which the subject process is conducted. This temperature range may conveniently be in between 0 °C to 250 °C, but preferably the subject process is conducted in the range of from 10 °C to 200 °C, and yet more preferably in the range of 15 °C to 150 °C, and again more preferably in the 15 range of from 18 °C to 130 °C. The rotation about the bond is considered hindered, if no rotation is measurable in the liquid state or in solution, i.e. at a temperature range where the ligand is liquid if determined under neat 20 conditions, or in solution of a suitable solvent.

Accordingly, the rotation about the bond C¹-C² of the bidentate ligand is hindered or restricted in solution or liquid state of the ligand. Preferably, the rotation about the bond C¹-C² of the bidentate ligand is 25 preferably hindered or restricted in the temperature range of the subject process.

The bridging group R^b comprises a chain of 2 optionally substituted carbon atoms C¹ and C². These carbon atoms C¹ and C² form the direct bridge between 30 R¹R²P¹-R^a_m- and -R^c_n-P²R⁵R⁶, so that the phosphorus atoms P¹ and P² and the optionally substituted methylene groups

R^a_m and R^c_n are connected via the bridging group C^1-C^2 to form the diphosphine ligands (b).

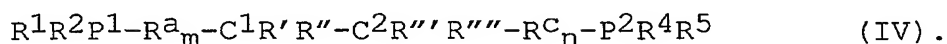
Although many different restricted conformations are possible for the subject ligands, a particular dihedral angle was found to be of high importance for the activity of the catalyst system. A dihedral angle is generally defined as the angle formed by two intersecting planes. The dihedral angle according to the subject process is the angle formed by the plane occupied by the three atom sequence composed of the three atoms C^2 , C^1 and the atom directly bonded to C^1 in direction of P^1 , and the plane occupied by the three atom sequence C^1 , C^2 and the atom directly bonded to C^2 in direction of P^2 is in the range of from 0 to 120°, of the four atom sequence (atom directly bonded to C^1 in direction of P^1)- C^1 - C^2 -(atom directly bonded to C^2 in the direction of P^2). "In the direction of P^1 or P^2 " herein has the meaning that the relevant atom is situated in that part of the ligand chain that connects C^1 and P^1 , or C^2 and P^2 , respectively.

For instance, in the case that m and n are equal to 1, the dihedral angle is the angle between the plane occupied by the three atom sequence $R^a-C^1-C^2$ of the four atom sequence $R^a-C^1-C^2-R^c$ and the other three atoms $C^1-C^2-R^c$ of the four atom sequence $R^a-C^1-C^2-R^c$. Each plane is understood to run through the central points of the respective atoms. In the case that m and n of formula (III) should equal 0, the four atom sequence would accordingly be $P^1-C^1-C^2-P^2$, and the two planes would be defined as $P^1-C^1-C^2$ and $C^1-C^2-P^2$.

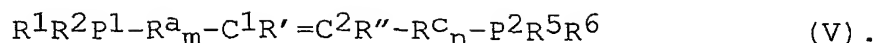
In the ligands according to the subject process, the dihedral angle as defined above is ranging from 0° to 120°. Since a higher catalytic activity of the catalyst system is thereby obtainable, the dihedral angle preferably is in the range of from 0° to 70, yet more preferably in the range of from 0° to 15°, and most preferably in the range of from 0° to 5°.

Without wishing to be bound to any particular theory, it is believed that ligands allowing rotation about the bond C¹-C² are less able to form a conformationally stable bidentate complex with the palladium centre.

The bond formed between C¹ and C² may be a saturated or an unsaturated bond as occurring in ethylenically unsaturated or aromatic compounds. In the case of a saturated bond connecting C¹ and C², R^b can be expressed by C¹R'R"-C²R'''R''', and the bidentate diphosphine ligand according to the present invention is thus suitably characterised by formula IV



In this embodiment, R' and R'', and R''' and R'''' represent hydrogen or the same or different optionally substituted organic group, provided that only one of R' and R'', and only one of R''' and R'''' is hydrogen. If C¹ and C² are connected by an ethylenically unsaturated double bond, C¹ and C² also cannot rotate freely. In this case, R can be expressed by C¹R'=C²R'', and the bidentate diphosphine ligand according to the present invention is thus suitably characterised by formula (V):



If the bond between C¹ and C² is an ethylenically unsaturated bond, the ligand chain connecting P¹ and P² via C¹ and C² may in principle exist in two isomeric forms, a trans-configuration, and a cis-configuration. According to the above definition, in the trans-configuration the dihedral angle is about 180°, whereas in the cis-configuration, the dihedral angle is about 0°.

The substituents R' to R''' in formula IV or V can themselves be independent substituents, thus only connected to each other via the carbon atoms C¹ and C², or preferably have at least one further connection. The substituents may further comprise carbon atoms and/or heteroatoms.

The restriction of the free rotation may conveniently be achieved by the bridging group C¹-C² forming part of a molecular structure that impedes rotation about the bond C¹-C² at ambient temperature, and more preferably at a temperature range from 0 to 250 °C, and preferably from 15 to 150 °C. This molecular structure may conveniently be for instance a) an ethylenically unsaturated double bond, wherein the rotation is impeded by the energetically advantageous overlap of π-bonds, and/or b) a cyclic hydrocarbyl structure, in which the rotation is restricted due to the steric interaction of substituents R' to R''', or due to steric strain induced by a cyclic structure formed by R' to R''' together, or by combination of the above factors, such as in aromatic or non-aromatic cyclic structures. Conformational stability and hence rigidity may also c) be achieved if the nature of the substituents R' and R'', and/or R''' and R''' is such that even if not connected to each other they impede

rotation about the bond C¹-C², for instance by strong steric interactions. To this goal, yet more preferably, none of R' to R''' in formula IV or V represent hydrogen.

R^b preferably is a cyclic hydrocarbyl structure that is optionally substituted by heteroatoms, yet more preferably an aliphatic or aromatic hydrocarbyl structure. This structure may be part of an optionally further substituted saturated or unsaturated polycyclic structure, which also optionally may contain heteroatoms such as nitrogen, sulphur, silicon or oxygen atoms.

Suitable structures R^b include for instance substituted cyclohexyl, cyclohexenyl, cyclohexadienyl, substituted cyclopentane, cyclopentenyl or cyclopentadienyl structures, all of which may optionally contain heteroatoms such as nitrogen, sulphur, silicon or oxygen atoms, with the proviso that the rotation about the bond C¹-C² is restricted, that the dihedral angle is in the range of from 0° to 120°, and that there is preferably no rotation about the bond formed by C¹ and C² induced by conformational changes, as for instance in highly restrained acetal structures such as 2, 2-dimethyl-1, 3-dioxolane. In one particularly preferred embodiment, R^b represents a divalent polycyclic hydrocarbyl ring structure. Such polycyclic groups are particularly preferred due to the high conformational stability and hence high restriction against free rotation about the bond between C¹ and C². Examples of such particularly preferred hydrocarbyl groups include norbornyl, norbornadienyl, isonobornyl, dicyclopentadienyl, octahydro-4, 7-methano-1H-indenemethanyl, α- and β-pinyl, and 1,

8-cineolyl, all of which may optionally be substituted, or contain heteroatoms as defined above.

In case that the bidentate ligand has chiral centres, it may be in any R,R-, S,S- or R,S-meso form, or mixtures thereof. Both meso forms and racemic mixtures can be employed, provided that the dihedral angle is in the range of from 0 to 120°.

In the diphosphine of formula II, R^b preferably represents an optionally substituted divalent aromatic group which is linked to the phosphorus atoms via the groups R^a and R^c. Such an aromatic cyclic structure is preferred due to its rigidity, and to a dihedral angle being generally in the range of 0 to 5 °.

The aromatic group can be a monocyclic group, such as for example a phenyl group or a polycyclic group, such as for example a naphthyl, anthryl or indyl group.

Preferably, the aromatic group R contains only carbon atoms, but R^b can also represent an aromatic group wherein a carbon chain is interrupted by one or more hetero atoms, such as nitrogen, sulphur or oxygen atom in for example a pyridine, pyrrole, furan, thiophene, oxazole or thiazole group. Most preferably the aromatic group R^b represents a phenyl group or naphthylene group.

Optionally the aromatic group R^b is substituted. Suitable substituents include groups containing hetero-atoms such as halides, sulphur, phosphorus, oxygen and nitrogen.

Examples of such groups include chloride, bromide, iodide and groups of the general formula -O-H, -O-X, -CO-X, -CO-O-X, -S-H, -S-X, -CO-S-X, -NH₂, -NHX, -NO₂, -CN, -CO-NH₂, -CO-NHX, -CO-NX₂, -CI₃ or -CF₃, in which X independently represents alkyl groups having from 1 to 4 carbon atoms like methyl, ethyl, propyl, isopropyl and n-butyl. When

the aromatic group R^b is substituted it is preferably substituted with one or more aryl, alkyl or cycloalkyl groups, preferably having from 1 to 10 carbon atoms. Suitable groups include methyl, ethyl, trimethyl, iso-propyl, tetramethyl and iso-butyl, phenyl and cyclohexyl.

Most preferably, however, the aromatic group R^b is non-substituted and only linked to the groups R^a and R^c which connect it with the phosphorus atoms. yet more preferably, R^a and R^c are alkylene groups, and thus connected at adjacent positions, for example the 1 and 2 positions, of the aromatic group.

The symbols m and n in formula III, IV and V independently may represent a natural number in the range of from 0 to 3. If the m and n are 0, then the phosphorus atoms P^1 and P^2 are directly connected to bridge formed by the carbon atoms C^1 and C^2 . If one of m or n equals 0, then either C^1 or C^2 will be directly connected to P^1 or P^2 . Without wishing to be bound to any particular theory, it is believed that the effect resulting from the particular arrangement of the central bridge formed by C^1 and C^2 on the phosphorus atoms, and hence on the catalyst complex, will be diluted by the presence of a larger number of groups R^a and/or R^c . Also, it is believed that if both m and n equal 0, the distance between the phosphorus atoms may be rather short, such that the ligand binds less strongly to the palladium centre atom of the catalyst complex. Accordingly, due to generally good catalyst activity found with such ligands, m preferably equals 0 or 1, whereas n preferably is in the range of from 1 to 3, more preferably from 1 to 2 and most preferably 1.

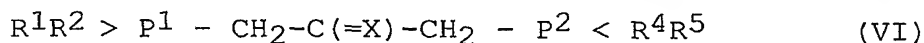
If m and/or n have a value above 1, then several optionally substituted groups R^a and R^c connect P^1 and P^2 to R^b . R^a and/or R^c . These different may then be the same or individually different groups. Hence, R^a and/or R^c preferably are lower alkylene groups (by lower alkylene groups is understood alkylene groups comprising from 1 to 4 carbon atoms). These alkylene groups can be substituted, for example with alkyl groups or heteroatoms, or non-substituted, and may for instance represent methylene, ethylene, trimethylene, isopropylene, tetramethylene, iso-butylene and tert-butylene, or may represent methyloxy, ethyloxy and similar groups. Most preferably, at least one of R^a and/or R^c is a methylene group.

Particularly suitable aromatic groups for R^b include aryl groups such as disubstituted phenyl or naphthyl groups, and substituted alkyl phenyl groups such as tolyl and xylyl groups. Preferred due to the high stability of the obtained ligands and catalysts, are tolyl and xylyl groups, wherein the methylene substituent or methylene substituents at the aromatic ring serve as groups R^a and/or R^c . Most preferably, C^1 and C^2 are part of an aromatic ring, whereas at least one of R^a and/or R^c represent methylene groups attached to adjacent atoms C^1 and C^2 in the aromatic ring.

Accordingly, an especially preferred ligand family according to the subject invention is that wherein C^1 and C^2 are part of a phenyl ring; m is 0 or 1; n is 1, and R^a and R^c are methylene groups. In yet another especially preferred ligand family due to easy synthetic accessibility, m and n equal 1. Accordingly, such ligands

based on the 1, 2-di(phosphinomethyl)benzene or
 1-phosphino-2-(phosphino methyl)-benzene groups are
 particularly suited for the subject process due to the
 high rigidity of the aromatic backbone, easy synthetic
 5 availability, and due to the very good results obtained
 with the derived catalyst system. Such a ligand is for
 instance 1-(1,3,5,7-tetramethyl-1,3, 5-trimethyl-6,9,
 10-trioxa-2-phosphatricyclo[3.3.1.1^{3,7}]^{3,7}decyl-2-(di-
 tert-butylphosphinomethyl)benzene.

10 Very good results were also obtained with
 unsymmetrical bidentate diphosphine ligand of formula VI,



wherein X represents an oxygen atom, a sulphur atom, an
 amino group, or a optionally substituted alkylene group,
 15 P¹ and P² represent phosphorus atoms; R¹, R², R⁴ and R⁵
 each individually, or R¹ and R² jointly, and/or R⁴ and R⁵
 jointly represent organic groups that are covalently
 linked to the phosphorus; and wherein R¹ to R⁴ are chosen
 in such way, that the phosphino group R¹R² > P¹ differs
 20 from the phosphino group P² < R⁴R⁵. In these ligands, R^a
 equals CH₂, m is 1, R^b equals C=CH₂, and R^c equals CH₂.

In these ligands, the divalent organic bridging group R³
 is represented by the tertiary alkylene carbon atom
 C(=X). X preferably is an optionally substituted alkylene
 25 group, more preferably a methylene group, so that R³
 equals -C(=CH₂)-. In these ligands, rotation about this
 carbon atom is also highly restricted due to the
 stiffness of the alkylene structure, and the dihedral
 angle formed by P¹-CH₂-C(=CH₂)- and C(=CH₂)-CH₂-P² is in
 30 the range of from 0° to 120°.

Preferably, these ligands were employed in a catalyst composition comprising the unsymmetrical bidentate diphosphine ligand and a source of palladium, platinum, or rhodium. The latter was found effective in particular for hydroformulation reaction, i.e. carbonylation reactions in the presence of a source of hydride, such as molecular hydrogen, in particular in the presence of a source of an anion.

The subject process permits to convert non-conjugated ethylenically and acetylenically unsaturated compounds to the corresponding products with very high turnover numbers due to the high catalyst activity. Suitable ethylenically or acetylenically unsaturated compounds are ethylenically or acetylenically unsaturated compounds having from 2 to 50 carbon atoms per molecule, or mixtures thereof. Suitable ethylenically and acetylenically unsaturated compounds may have one or more isolated unsaturated bonds per molecule. Preferred are compounds having from 2 to 20 carbon atoms, or mixtures thereof, yet more preferred are compounds having at most 18 carbon atoms, yet more at most 16 carbon atoms, again more preferred compounds have at most 10 carbon atoms. The ethylenically or acetylenically unsaturated compound may further comprise functional groups or heteroatoms, such as nitrogen, sulphur or oxide. Examples include carboxylic acids, esters or nitriles as functional groups. In a preferred embodiment, the ethylenically or acetylenically unsaturated compound is an olefin or a mixture of olefins. Such olefins can be converted by reaction with carbon monoxide and a co-reactant with a high regioselectivity towards the linear carbonylation product. Suitable ethylenically or acetylenically unsaturated compounds include acetylene, methyl

acetylene, propyl acetylene, ethylene, propylene, butylene, isobutylene, pentene, pentene nitriles, methyl 3-pentenoates, 2- and 3-pentenoic acid. Another preferred feedstock are mixtures of olefins, or mixtures of olefins and saturated compounds obtained for instance in a Fischer-Tropsch synthesis step. In such a synthesis, usually a mixture of olefins and saturated hydrocarbons is obtained. The content of olefins in this mixture may vary, and may be increased by starting of with the synthesis product of an iron-catalyzed Fischer-Tropsch synthesis step. An even higher olefinic content of the mixture may be obtained by catalytic dehydrogenation or thermal cracking in the presence of steam. The subject process in particular suited for reacting such mixtures, as it selectively reacts with olefinic or acetylenically non-conjugated compounds in such a mixture without affecting the other saturated compounds, and hence allows an easy separation of the carbonylated products from the saturated hydrocarbons. According to the subject process, ethylenically or acetylenically unsaturated compounds may be converted into very different chemical compounds, such as aldehydes, esters, acids or acid anhydride, thioesters, amides, and alcohols or others. The specific nature of the product is depending on the nature of the co-reactant having a mobile hydrogen atom. The co-reactant according to the present invention may thus be any compound having a mobile hydrogen atom, and which is capable of reacting as a nucleophile with the diene under catalysis. The nature of the co-reactant largely determines the type of product formed. A suitable co-reactant is water, a carboxylic acid, alcohol, ammonia or an amine, a thiol, or a combination thereof. In other words, the catalyst systems may be used in

"hydroformylation reactions", "hydrocarboxylation reactions", "hydroesterification reactions" or "hydroamidation reactions".

Inasmuch as the co-reactant is water, the product
5 obtained will be a carboxylic acid. Anhydrides are obtained inasmuch as the co-reactant is a carboxylic acid. For an alcohol co-reactant, the product of the carbonylation is an ester. Similarly, the use of ammonia (NH₃) or a primary or secondary amine RNH₂ or R'R''NH will
10 produce an amide, whereas the use of a thiol RSH will produce a thioester. In the above-defined co-reactants, R, R' and/or R'' represent optionally heteroatom-substituted organic radicals, preferably alkyl, alkenyl or aryl radicals. When ammonia or amines are employed, a
15 small portion of these co-reactants will react with acids present under formation of an amide and water. Hence, in the case of ammonia or amine-co-reactants, there is always water present. Preferred alcohol co-reactants are alkanols with 1 to 20, more preferably with 1 to 6 carbon
20 atoms per molecule, and alkanediols with 2-20, more preferably 2 to 6 carbon atoms per molecule. The alkanols can be aliphatic, cycloaliphatic or aromatic. Suitable alkanols in the process of the invention include methanol, ethanol, ethanediol, n-propanol, 1,3-
25 propanediol, iso-propanol, 1-butanol, 2-butanol (sec-butanol), 2-methyl-1-propanol (isobutanol), 2-methyl-2-propanol (tert-butanol), 1-pentanol, 2-pentanol, 3-pentanol, 2-methyl-1-butanol, 3-methyl-1-butanol (isoamyl alcohol), 2-methyl-2-butanol (tert-amyl alcohol), 1-
30 hexanol, 2-hexanol, 4-methyl-2-pentanol, 3,3-dimethyl-2-butanol, 1-heptanol, 1-octanol, 1-nonanol, 1-decanol, 1,2-ethylene glycol and 1,3-propylene glycol, of which methanol is the most preferred due to the high turn over

achievable and due to the particular usefulness of the obtained products. Preferred amines have from 1 to 20, more preferably 1 to 6 carbon atoms per molecule, and diamines have from 2-20, more preferably 2 to 6 carbon atoms per molecule. The amines can be aliphatic, cycloaliphatic or aromatic. More preferred due to the high turnovers achieved are ammonia and primary amines.

Thiol co-reactants can be aliphatic, cycloaliphatic or aromatic. Preferred thiol co-reactants are aliphatic thiols with 1 to 20, more preferably with 1 to 6 carbon atoms per molecule, and aliphatic dithiols with 2-20, more preferably 2 to 6 carbon atoms per molecule.

The subject carbonylation reaction is also known as a hydroformylation reaction in the case that the co-reactant is a hydride source, such as molecular hydrogen. This can be achieved by supplying carbon monoxide and hydrogen to the reaction in equimolar or non-equimolar ratios, e.g. in a ratio within the range of 5:1 to 1:5, preferably 3:1 to 1:3. Preferably they are supplied in a ratio within the range of 2:1 to 1:2, and will result in alcohol mixtures. In particular when Fischer-Tropsch product streams are used as feedstocks, the obtained product mixtures can be suitably used for sanitary or other related applications.

The carbonylation can be suitably carried out at moderate reaction conditions. Hence temperatures in the range of 50 to 200 °C are recommended, preferred temperatures being in the range of 70 to 160 °C. Reaction pressures in the range of 5 to 100 bar are preferred, lower or higher pressures may be selected, but are not considered particularly advantageous. Moreover, higher pressures require special equipment provisions.

In the process of the invention, the unsaturated starting material and the formed product as well as the co-reactant may act as reaction solvent. Hence, the use of a separate solvent might not be necessary.

5 Conveniently, however, the carbonylation reaction may be carried out in the additional presence of a solvent, in particular in the start-up phase of the reaction. As such, saturated hydrocarbons, e.g. paraffins and isoalkanes are recommended and furthermore alcohols, the
10 saturated hydrocarbons and alcohols preferably having from 1 to 10 carbon atoms per molecule, such as methanol, butanol, 2-ethylhexan-1-ol, nonan-1-ol, or in general terms the alcohols formed as carbonylation product; ethers such as 2,5,8-trioxanonane (diglyme), diethylether
15 and anisole, and ketones such as methylethylketone. Solvents, comprising or substantially consisting of sulphones are also preferred. Sulphones are in particular preferred, for example dialkylsulphones such as dimethylsulphone and diethylsulphone and cyclic sulphones, such
20 as sulfolane, 2-methylsulfolane and 2-methyl-4-ethylsulfolane.

The ratio of moles of bidentate diphosphine, i.e. catalyst component (b), per mole atom of palladium cations, i.e. catalyst component (a), ranges from 0.5 to
25 10, preferably from 0.8 to 8, and yet more preferably from 1 to 5. The ratio of moles of bidentate diphosphine, i.e. catalyst component (b), per mole atom of palladium, i.e. catalyst component (a), is not critical. Preferably it ranges from 0.1 to 100, more preferably from 0.5 to
30 10.

A more preferred catalyst the active species is based on an equimolar amount of bidentate diphosphine ligand per mole palladium. Thus, the molar amount of bidentate

diphosphine ligand per mole palladium is preferably in the range of 1 to 3, more preferably in the range of 1 to 2, and yet more preferably in the range of 1 to 1.5. In the presence of oxygen, slightly higher amounts may be beneficial.

Preferably, the catalyst system or composition further preferably comprises a source of an anion. A suitable source of anion may be a strong, or a weak coordinating anion. Various anions may be used as counter-ion to the metal cation. Examples thereof include anions that are the conjugated base of acids having a pKa (measured at 18 °C in water) of less than 6, preferably less than 4. The anions derived from these acids do not or only weakly co-ordinate with the metal cation, by which is meant that little or no covalent interaction occurs between the anion and the cation. Catalysts based on these anions exhibit a good activity.

Suitable anions include anions derived from Bronsted acids, such as from phosphoric acid and sulphuric acid, and in particular from sulphonic acids and (halogenated) carboxylic acids, such as trifluoroacetic acid, 2,6-dichlorobenzoic acid, and 2,6-bis(trifluoromethyl)benzoic acid or trifluoroacetic acid, etc. Anions derived from sulphonic acids are particularly preferred, for example methanesulphonic acid, trifluoromethanesulphonic acid, tert-butanesulphonic acid, p-toluenesulphonic acid and 2,4, 6-trimethylbenzenesulphonic acid. Also complex anions are suitable, such as the anions generated by a combination of a Lewis acid such as BF_3 , $\text{B}(\text{C}_6\text{F}_5)_3$, AlCl_3 , SnF_2 , $\text{Sn}(\text{CF}_3\text{SO}_3)_2$, SnCl_2 or GeCl_2 , with a protic acid, preferably having a pKa of less than 5, such as a sulphonic acid, e.g. $\text{CF}_3\text{SO}_3\text{H}$ or $\text{CH}_3\text{SO}_3\text{H}$ or a

hydrohalogenic acid such as HF or HCl, or a combination of a Lewis acid with an alcohol. Examples of such complex anions are BF_2^- , SnCl_3^- , $[\text{SnCl}_2.\text{CF}_3\text{SO}_3]^-$ and PF_6^- .

Preferably, the source of anions is a carboxylic acid, which can serve both as promoter component, as well as solvent for the reaction. Again more preferably, the source of anions is an acid having a pK_a below 3.6 (measured in aqueous solution at 18 °C), and yet more preferably catalyst component (c) is an acid having a pK_a below 3.0, and yet more preferably a pK_a of above 2.0.

Examples of suitable acids include acetic acid, propionic acid, butyric acid, pentanoic acid, pentenoic acid and nonanoic acid. Very conveniently the acid corresponding to the desired product of the reaction can be used as the catalyst component (c). Catalyst component (c) can also be an ion exchanging resin containing carboxylic acid groups. This advantageously simplifies the purification of the product mixture.

In the case that the co-reactant is an amine or ammonia, and the source of an anion of the catalyst system is an acid, preferably the amount of ammonia or amine is less than stoichiometric based on the amine functionality. Inadvertently, when the co-reactant is ammonia, and to a lesser extent a primary amine, a small amount of the acid present will react to an amide under liberation of water. Hence, there is also always a small amount of acid formed from the unsaturated compound, carbon monoxide and the water, which in turn replaces acid converted to amide by the direct reaction as described above.

The molar ratio of the source of anions, and palladium, platinum, or rhodium is not critical. However,

it suitably is between 2:1 and 10^7 :1 and more preferably between 10^2 :1 and 10^6 :1, yet more preferably between 10^2 :1 and 10^5 :1, and most preferably between 10^2 :1 and 10^4 :1 due to the enhanced activity of the catalyst system. Accordingly, if a co-reactant should react with the acid serving as source of anions, then the amount of the acid to co-reactant should be chosen such that a suitable amount of free acid is present. Generally, a large surplus of acid over the co-reactant is preferred due to the enhanced reaction rates.

The quantity in which the complete catalyst system is used is not critical and may vary within wide limits. Usually amounts in the range of 10^{-8} to 10^{-1} , preferably in the range of 10^{-7} to 10^{-2} mole atom of palladium per mole of olefin are used, preferably in the range of 10^{-5} to 10^{-2} gram atom per mole. The process may optionally be carried out in the presence of a solvent, however preferably the acid is used as solvent and as promoter.

The carbonylation reaction according to the present invention is carried out at moderate temperatures and pressures. Suitable reaction temperatures are in the range of 0-250 °C, more preferably in the range of 50-200 °C, yet more preferably in the range of from 80-150 °C.

The reaction pressure is usually at least atmospheric. Suitable pressures are in the range of 0.1 to 15 MPa (1 to 150 bar), preferably in the range of 0.5 to 8.5 MPa (5 to 85 bar). Carbon monoxide partial pressures in the range of 0.1 to 8 MPa (1 to 80 bar) are preferred, the upper range of 4 to 8 MPa being more preferred. Higher pressures require special equipment provisions.

The carbonylation reaction is further conveniently carried out at a temperature in the range of 30 to 200 °C, preferred temperatures being in the range of 50 to 180 °C.

In the process according to the present invention, the carbon monoxide can be used in its pure form or diluted with an inert gas such as nitrogen, carbon dioxide or noble gases such as argon, or co-reactant gases such as ammonia.

The invention will be illustrated by the following non-limiting examples.

Example 1: Preparation of 9-(2-{[di(tert-butyl)phosphino]methyl}-2-propenyl)-9-phosphabicyclo[3.3.1]nonane
9-phosphabicyclo[3.3.1]nonane

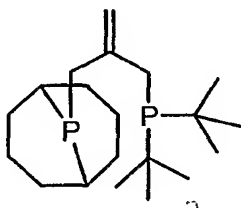


Figure 1: 9-(2-{[di(tert-butyl)phosphino]methyl}-2-propenyl)-

A mixture of di-*tert*-butylphosphine (5 g, 34.1 mmol) and 3-chloro-2-chloromethyl-1-propene (14,04 g, 69 mmol) was dissolved in degassed acetonitrile (70 ml) and heated to 60 °C for 16 hours. The solvent was removed in vacuo and the product was suspended in hexane (50 ml). After filtration the product was washed with hexane (2 times 40 ml). The white solid salt was neutralized in a mixture of toluene (50 ml) and water (30 ml) with HNEt₂ (5 ml) under reflux conditions. The toluene layer was separated, washed with water (2 times 30 ml) and removed in vacuo, to yield 3.14 g (40%) of di(*tert*-butyl)[2-(chloromethyl)-2-propenyl]phosphine.

The product could be characterized by showing a distinct resonance signals in ^{31}P NMR at $\delta = + 21.7$ ppm.

2.82 g (20 mmol) of the 9-phoshabicyclo[3.3.1]nonane was dissolved in THF (90 ml). To this solution, a 1.6 M solution of n-BuLi in hexane (12.5 ml, 20 mmol) at -50°C was added. The mixture was allowed to warm to 0°C for fifteen minutes and then cooled to -70°C , after which and a solution of 4.7 g (20 mmol) di(tert-butyl)[2-(chloromethyl)-2-propenyl]-phosphine obtained as described above in 10 ml THF was added. The mixture was allowed to warm to ambient temperature, and was subsequently quenched with water (5 ml). The solvents were removed in vacuo and toluene (60 ml) was added. The mixture was washed twice with water (40 ml and 20 ml) and the toluene was removed in vacuo giving 6.71 g of a red/orange oil. The product was crystallized twice from (boiling) methanol, giving a white solid (2.83 g, 42%). The product could be characterized by showing two distinct resonance signals in ^{31}P NMR at $+ 22.1$ ppm and -36.5 ppm.

Example 2: Preparation of 8-(2-([di(tert-butyl)phosphino]methyl)-2-propenyl)-1,3,5,7-tetramethyl-2,4,6-trioxa-8-phosphatricyclo[3.3.1.1.{3,7}]decane

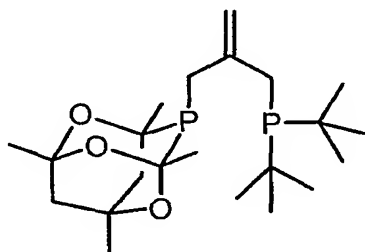


Figure 2: 8-(2-([di(tert-butyl)phosphino]methyl)-2-propenyl)-1,3,5,7-tetramethyl-2,4,6-trioxa-8-phosphatricyclo[3.3.1.1.{3,7}]decane

To a solution of HPCage* BH_3 (1.34 g, 5.83 mmol; P_{cage} = 1,3,5,7-tetramethyl-2,4,6-trioxa-8-phosphatricyclo[3.3.1.1.1-{3,7}]decane) in THF (20 ml) was added a 1.6 M solution of *n*-BuLi in hexane (3.64 ml, 5.83 mmol) at -78 °C. The mixture was allowed to warm to 0 °C for fifteen minutes and cooled to -78 °C again, after which a solution of di(*tert*-butyl)[2-(chloromethyl)-2-propenyl]-phosphine as described under (1.37 g, 5.83 mmol) in THF (5 ml) was added. The reaction mixture was allowed to warm to RT. HNEt₂ (10 ml) was added and the reaction mixture was refluxed for a few hours, until NMR-measurements showed that the deprotection was complete). Subsequently, the solvents were removed in vacuo, and the solid product was recrystallized from methanol (50 ml) upon gradual cooling. The isolated yield was 1.5 g (62%), and the product could be characterized by showing two distinct resonance signals in ³¹P NMR at + 20.3 ppm (P-*tert*.Bu) and - 34.3 ppm (P-Cage).

Example 3: Preparation of 8-(2-[[di(*tert*-butyl)phosphino]methyl]phenyl)-1,3,5,7-tetramethyl-2,4,6-trioxa-8-phosphatricyclo[3.3.1.1{3,7}]decane

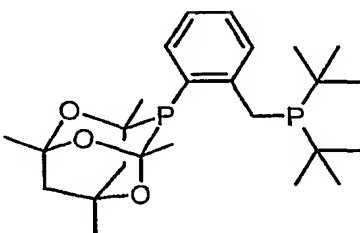


Figure 3: 8-(2-[[di(*tert*-butyl)phosphino]methyl]phenyl)-1,3,5,7-tetramethyl-2,4,6-trioxa-8-phosphatricyclo[3.3.1.1{3,7}]decane

8.25 g (33 mmol) 2-bromobenzylbromide and 5 g (34.2 mmol) di-*tert*-butyl phosphine in 40 ml degassed acetonitrile were measured into a 100 ml glass reactor

under an inert atmosphere, and then stirred for a period of 12 hours at ambient temperature. The acetonitrile was then removed in vacuo and 30 ml degassed toluene, 30 ml degassed water and 7.5 ml triethylamine were added. To this mixture 10 ml ethanol was added to improve phase separation. Upon phase separation, the upper layer containing the toluene was separated and evaporated to dryness. The remainder was 9 g (28.6 mmol, 87%) of (2-bromobenzyl)(di-*tert*-butyl)phosphine as a light yellow oil exhibiting a resonance peak in ^{31}P NMR at +34.16 ppm.

2.5 g (7.9 mmol) of the thus obtained 2-bromobenzyl (di-*tert*-butyl)phosphine, 2.24 g DABCO (20 mmol), 1.94 g 1,3,5-trimethyl-4,6,9-trioxa-2-phosphatricyclo-[3.3.1.1^{3,7}]-decane (9 mmol) and 0.23 g $\text{Pd}(\text{PPh}_3)_4$ (0.2 mmol) in 10 ml toluene were added into a 250 ml glass vessel under inert atmosphere, and the content of the vessel was heated to 140 °C under stirring for 12 hours. The mixture was then allowed to cool to 100 °C, and was then filtered. The filtrate was cooled to room temperature, then 30 ml of methanol added were added and the mixture cooled for a period for 12 hours to -35 °C, 8-(2-([di(*tert*-butyl)-phosphino]-methyl)phenyl)-1,3,5,7-tetramethyl-2,4,6-trioxa-8-phosphatricyclo[3.3.1.1^{3,7}]-decane was isolated as yellow crystals (2.2 g, 4.9 mmol, 62%), and could be characterized by showing two distinct resonance signals in ^{31}P NMR at +38.08 and -38.96 ppm.

Examples 4 to 6 and Comparative Examples A, B and C:
Conversion of 1-octene with methanol

A 250 ml magnetically stirred autoclave, made of HASTELLOY C (HASTELLOY C is a trademark), was successively charged with 40 ml of methanol, 20 ml 1-octene, and 0.5 mmol of methanesulfonic acid, and a

catalyst solution comprising 0.1 mmol palladium acetate and 0.2 mmol of the respective ligand in 10 ml of methanol. The autoclave was closed, evacuated, and pressurized with 3 MPa carbon monoxide. The reactor was then heated to 70°C, and maintained at this temperature for 10 hours. Finally the autoclave was cooled, depressurized, and the reaction mixture was analysed with GLC. The turn over frequency is defined as amount of converted products, and given in mol olefin per mol Pd per hour for the batch operation as presented in Table I.

Examples 7 and 8: Conversion of ethylene with methanol

A 250 ml magnetically stirred autoclave, made of HASTELLOY C (HASTELLOY C is a trademark), was successively charged with 45 ml of methanol, and 0.5 mmol of methanesulfonic acid. The autoclave was closed, evacuated, and pressurized with 3 MPa carbon monoxide and 20 MPa ethylene. The reactor was then heated to 85 °C, and a catalyst solution comprising 0.05 mmol palladium acetate and 0.1 mmol of the respective ligand in 5 ml of methanol was injected into the autoclave, thereby raising the carbon monoxide pressure by an additional 5 bar. The autoclave was maintained at 85 °C for 10 hours. Finally the autoclave was cooled, depressurized, and the reaction mixture was analysed with GLC. The turn over frequency is defined as amount of converted products, and given in mol olefin per mol Pd per hour for the batch operation as presented in Table I.

In Example 4 according to the invention, the ligand was 9-(2-([di(tert-butyl)phosphino]methyl)-2-propenyl)-9-phosphabicyclo[3.3.1]nonane prepared in Example 1. In Examples 5 and 7 according to the invention, the ligand was 8-(2-([di(tert-butyl)phosphino]methyl)-2-propenyl)-1,3,5, 7-tetramethyl-2,4,6-trioxa-8-

phosphatricyclo[3.3.1.1.3,7]decane prepared in Example 2.
 In Examples 6 and 8 according to the invention, the
 ligand was 8-(2-{[di(*tert*-butyl)phosphino]methyl}phenyl)-
 1,3,5, 7-tetramethyl-2,4, 6-trioxa-8-phospha-
 5 tricyclo[3.3.1.1{3,7}]decane prepared in Example 3.

In Comparative Example A, the ligand was a
 symmetrical ligand with two 9-phosphabicyclo[3.3.1]nonane
 rings attached to an isobutene structure (9-[2-(9-
 phosphabicyclo[3.3.1]non-9-ylmethyl)-2-propenyl]-9-
 10 phosphabicyclo[3.3.1]nonane). In Comparative Example B,
 the ligand was the symmetrical bis-di-*tert*-butyl
 phosphino-analogue (di(*tert*-butyl)(2-{[di(*tert*-
 butyl)phosphino]methyl}-2-propenyl)phosphine). In
 Comparative Example C, the ligand was the symmetrical
 15 bis-trioxa-adamantyl-ligand (1,3,5,7-tetramethyl-8-{2-
 [1,3,5, 7-tetramethyl-2,4,6-trioxa-8-
 phosphatricyclo[3.3.1.1{3,7}]dec-8-yl)methyl]-2-
 propenyl}-2,4,6-trioxa-8-phospha-
 tricyclo[3.3.1.1{3,7}]decane). In Comparative Example D,
 20 the ligand was the symmetrical bis-trioxadamantyl-ligand
 ([1,3,5,7-tetramethyl-8-{2-[1,3,5, 7-tetramethyl-2,4,6-
 trioxa-8-phosphatricyclo[3.3.1.1{3,7}]dec-8-
 yl)methyl]benzyl}-2,4,6-trioxa-8-phosphatricyclo-
 [3.3.1.1{3,7}]decane). In Comparative Example E, the
 25 ligand was the symmetrical 1,2-bis(di-*tert*-
 butylphosphinomethyl)benzene ([1,2-phenylenebis-
 (methylene)]bis(di-*tert*-butylphosphine)).

The results are depicted in Table 1.

Table I

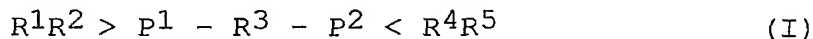
Example	Ligand	Substrate	Turn over Frequency [mol/mol Pd/hr]	Selectivity for linear product, [%]
Ex. 4	Ex. 1	1-Octene	1500	83
Ex. 5	Ex. 2	1-Octene	15000	87
Ex. 6	Ex. 3	1-Octene	20,000	96
Comp. A	see above	1-Octene	150	76
Comp. B	see above	1-Octene	250	67
Comp. C	see above	1-Octene	500	83
Comp. D	see above	1-Octene	50	67
Comp. E	see above	1-Octene	100	97
				Chemoselect. carbonylation product, [%]
Ex. 7	Ex.5	Ethene	20,000	99,9
Ex. 8	Ex.6	Ethene	60,000	99,9

From the above data, it is clear that the unsymmetrical ligands clearly outperform the symmetrical ligands in catalyst activity and selectivity for the desired linear products. Moreover, the use of two different kinds of substituents R^1 , R^2 , R^4 and R^5 , wherein R^1 and R^2 have a different electron-donating ability on P^1 as compared to R^4 and R^5 on P^2 , and wherein all substituents R^1 , R^2 , R^4 and R^5 are bonded to the phosphorus atoms via tertiary carbon atoms results in yet unprecedented catalyst activity and selectivity for the linear carbonylation products.

C L A I M S

1. Process for the carbonylation of unsaturated compounds by contacting the unsaturated compound with carbon monoxide in the presence of a catalyst system comprising:

- 5 (a) a source of palladium and/or platinum; and
(d) an unsymmetrical bidentate diphosphine ligand of formula I,



wherein P^1 and P^2 represent phosphorus atoms; R^3

10 represents a divalent organic bridging group; and R^1 , R^2 , R^4 and R^5 each individually, or R^1 and R^2 jointly, and/or R^4 and R^5 jointly represent organic groups that are covalently linked to the phosphorus; and wherein R^1 , R^2 , R^4 and R^5 are chosen in such way, that the phosphino group $R^1R^2 > P^1$ differs from the phosphino group $P^2 < R^4R^5$.

2. A process according to claim 1, wherein R^1 and R^2 have higher electron donor ability on P^1 than R^4 and R^5 on P^2 .

3. A process according to claim 1 or claim 2, wherein R^1 and R^2 independently represent the same or different optionally substituted organic group and which groups are solely connected to each other via the phosphorus atom P^1 ; and wherein R^4 and R^5 together represent an organic bivalent group linked to the phosphorus atom P^2 .

4. A process according to any one of claims 1 to 3, wherein R^1 , R^2 , R^4 and R^5 each represent an organic radical which is covalently connected to the respective phosphorus atom via a tertiary carbon atom.

5. A process according to any one of claims 1 to 4, wherein R^1 and R^2 together, or R^4 and R^5 together represent a 2-phospha-adamantane structure or a phosphinan-4-one structure.

6. A process according to any one of claims 1 to 5, wherein wherein R^3 represents a group



wherein R^a and R^b independently represent the same or different optionally substituted methylene groups;

R^b represents an organic group comprising a bivalent bridging group C^1-C^2 through which R^b is connected to R^a and R^c ; m and n independently represent a natural number in the range of from 0 to 4, and wherein the rotation about the bond between the carbon atoms C^1 and C^2 of the bridging group is restricted at a temperature in the range of from 0 °C to 250 °C, and wherein the dihedral angle between the plane occupied by the three atom sequence composed of C^1 , C^2 and the atom directly linked to C^1 in the direction of P^1 , and the plane occupied by the three atom sequence C^1 , C^2 and the atom directly linked to C^2 in the direction of P^2 , is in the range of from 0 to 120°.

7. A process according to claim 7, wherein R^b represents an aryl group, and wherein m is 0 or 1, n is 1, and the phosphorus atoms P^1 and P^2 are connected to R through the

respective methylene groups R^a or R^c or directly in adjacent positions on the aryl group.

8. A process according to claim 6 or 7, wherein R¹ and R² represent tertiary butyl groups, and R⁴ and R⁵ together represent a 2-phospha-adamantane structure or a phosphinan-4-one structure.

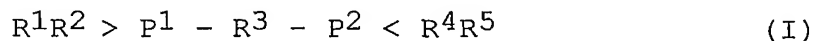
9. A process according to anyone of claims 1 to 8, wherein ethylene is converted to methyl propionate in presence of methanol.

10. A process according to anyone of claims 1 to 8, wherein pentenoic acid is converted to adipic acid.

11. A catalyst composition comprising:

(a) a source of palladium, and/or platinum; and

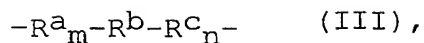
(b) an unsymmetrical bidentate diphosphine ligand of formula I,



wherein P¹ and P² represent phosphorus atoms; R³ represents a divalent organic bridging group; and R¹, R², R⁴ and R⁵ each individually, or R¹ and R² jointly, and/or R⁴ and R⁵ jointly represent organic groups that are covalently linked to the phosphorus; and wherein R¹, R², R⁴ and R⁵ are chosen in such way, that the phosphino group R¹R² > P¹ differs from the phosphino group P² < R⁴R⁵.

12. A catalyst composition according to claim 11, wherein R¹ and R² represent tertiary butyl groups, and R⁴ and R⁵ together represent a 2-phospha-adamantane structure or a phosphinan-4-one structure.

13. A catalyst composition according to claims 11 or 12, wherein R³ represents a group

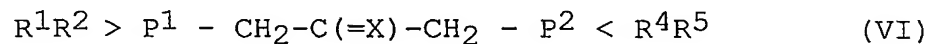


wherein R^a and R^b independently represent the same or different optionally substituted methylene groups;
 R^b represents an organic group comprising a bivalent
 5 bridging group C^1-C^2 through which R^b is connected to R^a and R^C ; m and n independently represent a natural number in the range of from 0 to 4, and wherein the rotation about the bond between the carbon atoms C^1 and C^2 of the bridging group is restricted at a temperature in the
 10 range of from 0 °C to 250 °C, and wherein the dihedral angle between the plane occupied by the three atom sequence composed of C^1 , C^2 and the atom directly linked to C^1 in the direction of P^1 , and the plane occupied by the three atom sequence C^1 , C^2 and the atom directly
 15 linked to C^2 in the direction of P^2 , is in the range of from 0 to 120°.

14. A catalyst composition according to claim 13, wherein
 R^b represents an aryl group, and wherein m is 0 or 1, n is 1, and the phosphorus atoms P^1 and P^2 are connected to
 20 R through the respective methylene groups R^a or R^C or directly in adjacent positions on the aryl group.

15. A catalyst composition according to any one of claims 11 to 14, further comprising a source of an anion.

16. An unsymmetrical bidentate diphosphine ligand of
 25 formula VI,



wherein X represents an oxygen atom, a sulphur atom, an amino group, or a optionally substituted alkylene group;
 wherein P^1 and P^2 represent phosphorus atoms; R^1 , R^2 , R^4
 30 and R^5 each individually, or R^1 and R^2 jointly, and/or R^4

and R⁵ jointly represent organic groups that are covalently linked to the phosphorus; and wherein R¹ to R⁴ are chosen in such way, that the phosphino group R¹R² > P¹ differs from the phosphino group P² < R⁴R⁵.

5 17. An unsymmetrical bidentate diphosphine ligand according to claim 16, wherein X represents a methylene group.

10 18. A catalyst composition comprising an unsymmetrical bidentate diphosphine ligand according to claim 16 or claim 17, and a source of palladium, platinum, or rhodium.

19. A catalyst composition according to claim 18, further comprising a source of an anion.

INTERNATIONAL SEARCH REPORT

International Application No
PCT/EP2005/050803

A. CLASSIFICATION OF SUBJECT MATTER

IPC 7 C07C67/38 C07C69/533 C07C51/14 C07C57/03 C07F9/6571
C07F9/6568 C07F9/28 B01J31/28 C07F9/6568 C07C67/38
C07C51/14

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

IPC 7 C07C B01J C07F

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

EPO-Internal, WPI Data, PAJ, CHEM ABS Data

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 1 121 367 A (E.I. DU PONT DE NEMOURS AND COMPANY) 8 August 2001 (2001-08-08) paragraph '0023!; examples 11-41	11, 13
A	DE 101 48 712 A1 (BASF AG) 17 April 2003 (2003-04-17) the whole document	1-19
A	US 2003/073868 A1 (HOGG GARRETT STEWART ET AL) 17 April 2003 (2003-04-17) the whole document	11-19
A	EP 0 273 489 A (SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ B.V) 6 July 1988 (1988-07-06) the whole document	1-19
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☒ Further documents are listed in the continuation of box C.

☒ Patent family members are listed in annex.

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- * & * document member of the same patent family

Date of the actual completion of the international search

23 June 2005

Date of mailing of the international search report

05/07/2005

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INTERNATIONAL SEARCH REPORT

International Application No
PCT/EP2005/050803

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT		
Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 5 495 041 A (SIELCKEN ET AL) 27 February 1996 (1996-02-27) the whole document	1-19
A	WO 02/064249 A (BASF AKTIENGESELLSCHAFT; SLANY, MICHAEL; SCHAEFER, MARTIN; ROEPER, MIC) 22 August 2002 (2002-08-22) the whole document	1-19
A	WO 01/68583 A (SHELL INTERNATIONALE RESEARCH MAATSCHAPPIJ B.V) 20 September 2001 (2001-09-20) the whole document	1-19

INTERNATIONAL SEARCH REPORT

International Application No
PCT/EP2005/050803

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
EP 1121367	A	08-08-2001	AT 239748 T	15-05-2003
			CA 2345640 A1	20-04-2000
			DE 69907712 D1	12-06-2003
			DE 69907712 T2	25-03-2004
			EP 1121367 A1	08-08-2001
			JP 2002527444 T	27-08-2002
			US 2002169340 A1	14-11-2002
			WO 0021970 A1	20-04-2000
			US 2003225226 A1	04-12-2003
			US 6137012 A	24-10-2000
			US 6350903 B1	26-02-2002
			US 2004225092 A1	11-11-2004
DE 10148712	A1	17-04-2003	BR 0212590 A	13-10-2004
			CA 2461920 A1	17-04-2003
			CN 1564825 A	12-01-2005
			WO 03031457 A1	17-04-2003
			EP 1434781 A1	07-07-2004
			JP 2005504850 T	17-02-2005
			MX PA04002504 A	31-05-2004
			US 2004249216 A1	09-12-2004
US 2003073868	A1	17-04-2003	US 2002143214 A1	03-10-2002
			US 2005124830 A1	09-06-2005
			US 2004138487 A1	15-07-2004
			BR 0200846 A	25-03-2003
			CA 2377069 A1	19-09-2002
			EP 1243591 A2	25-09-2002
			JP 2002338586 A	27-11-2002
			MX PA02001147 A	21-04-2004
EP 0273489	A	06-07-1988	AT 65489 T	15-08-1991
			AU 597451 B2	31-05-1990
			AU 8220187 A	16-06-1988
			CA 1292475 C	26-11-1991
			CN 87107325 A ,C	22-06-1988
			DE 3771686 D1	29-08-1991
			EP 0273489 A2	06-07-1988
			JP 2867137 B2	08-03-1999
			JP 63156745 A	29-06-1988
			KR 9601890 B1	06-02-1996
			US 5028734 A	02-07-1991
US 5495041	A	27-02-1996	CN 1141286 A	29-01-1997
			EP 0728733 A1	28-08-1996
			JP 8245509 A	24-09-1996
WO 02064249	A	22-08-2002	DE 10106348 A1	14-08-2002
			AT 278467 T	15-10-2004
			BR 0206982 A	03-08-2004
			CA 2437602 A1	22-08-2002
			DE 50201214 D1	11-11-2004
			WO 02064249 A2	22-08-2002
			EP 1360005 A2	12-11-2003
			JP 2004523544 T	05-08-2004
			MX PA03006490 A	15-10-2003
			US 2004059157 A1	25-03-2004

INTERNATIONAL SEARCH REPORT

Information on patent family members

International Application No

PCT/EP2005/050803

Patent document cited in search report		Publication date	Patent family member(s)	Publication date
WO 0168583	A	20-09-2001	AT 266617 T	15-05-2004
			AT 266628 T	15-05-2004
			AU 4246301 A	24-09-2001
			AU 5038001 A	08-10-2001
			BR 0109239 A	24-12-2002
			BR 0109242 A	24-12-2002
			CA 2403183 A1	04-10-2001
			CA 2403186 A1	20-09-2001
			CN 1420860 A	28-05-2003
			CN 1422248 A	04-06-2003
			DE 60103254 D1	17-06-2004
			DE 60103254 T2	25-05-2005
			DE 60103255 D1	17-06-2004
			DE 60103255 T2	25-05-2005
			WO 0172697 A2	04-10-2001
			WO 0168583 A2	20-09-2001
			EP 1263709 A2	11-12-2002
			EP 1263713 A2	11-12-2002
			JP 2003527365 T	16-09-2003
			JP 2003528849 T	30-09-2003
			MX PA02008927 A	25-04-2003
			MX PA02008928 A	25-04-2003
			US 2002045748 A1	18-04-2002
			US 2001044556 A1	22-11-2001
			US 2004230071 A1	18-11-2004